

## Group 5 / VB; Vanadium Group

Vanadium (V), Niobium (Nb) and tantalum (Ta) are important metals that are not all that rare.

- Tantalum has been seen as critical since it was linked to the ongoing war in the Congo, a source of the mineral group known as "coltan". Wrongly, tantalum was and is described as crucial for silicon microchip technology. It is actually important for so-called tantalum (or better tantalum oxide) capacitors, which are indeed crucial for electronics but not related to silicon technology.
- Vanadium, Niob and Tantal are important and relatively common elements; Hahnium is artificially produced. All elements are technically important ("Tantalkondensator"), especially as alloying components. Nb is (and is likely) a main component of technically used superconductors.
- Niobium and vanadium are important alloy elements for steel. Vanadium is also held responsible for the "water pattern" in wootz steel.

*Table of Basic Data*

| Name<br>(German)   | Vanadium<br><i>Vandium</i> | Niob<br><i>Niobium</i> | Tantal<br><i>Tantalum</i> |
|--|----------------------------|------------------------|---------------------------|
| Atomic number  | 23                         | 41                     | 73                        |
| Atomic mass [u]  | 50,9415                    | 92,90638               | 180,9479                  |
| Melting point [K]  | 2163                       | 2741                   | 3269                      |
| Melting point [°C]   | 1890                       | 2468                   | 2996                      |
| Melting point [°F]   | 3434                       | 4474                   | 5425                      |
| Boiling point [K]  | 3653                       | 5200                   | 5698                      |
| Density [g/cm³]  | 6,09                       | 8,58                   | 16,68                     |
| Ionization energy [eV]   | 6,74                       | 6,88                   | 7,89                      |
| Electro-negativity   | 1,5                        | 1,2                    | 1,4                       |
| Atomic radius [pm]   | 132,1                      | 142,9                  | 143                       |
| Ionic radius [pm]  | 59                         | 69                     | 64                        |
| Oxidation numbers  | 5 bis -3                   | 5 bis -1               | 5 bis -1                  |
| Lattice typ<br>Transformation temp. [°C]   | bcc<br>-                   | bcc<br>-               | bcc<br>-                  |
| Therm. expansion coefficient $\alpha$<br>Lattice constant [ $\text{\AA}$ ]<br>(a or c) | 3,03                       | 3,30                   | 3,30                      |
| Young's - Modulus [GPa]  | 127                        | 104                    | 175                       |
| [ $10^{-6} \text{K}^{-1}$ ]  | 7,8                        | 7,1                    | 3,6                       |

- In case of doubt all numbers are for room temperatures
- fcc = [face centered cubic](#); lattice const. = a  
bcc = [body centered cubic](#)

sc = [simple cubic](#)

hp = simple [hexagonal](#)

hcp = [hexagonal close packed](#); lattice constants a and c.

op = [simple orthorhombic](#), [monoclinic](#), [triclinic](#)

tp = [simple tetragonal](#)

dia = [diamond structure](#)

r = [trigonal](#) or rhomboedral trigonal